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$4\text{NH}_3(\text{g}) + 5\text{O}_2(\text{g}) \rightleftharpoons 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$ $K_c = 2.0 \times 10^{-2}$ Describe the effect, if any, of each of the following changes on the equilibrium concentration of nitrogen monoxide in a particular equilibrium mixture, giving a reason in each case. (i) Increasing the pressure, at constant temperature [2]

CHEMISTRY STANDARD LEVEL PAPER 2

- 2 - Section A Answer all questions. Write your answers in the boxes provided. 1. Ethanedioic acid is a diprotic acid. A student determined the value of x in the formula of hydrated ethanedioic acid, $\text{HOOC}-\text{COOH} \cdot x\text{H}_2\text{O}$, by titrating a known mass of the acid with a $0.100 \text{ mol dm}^{-3}$ solution of $\text{NaOH}(\text{aq})$. 0.795 g of ethanedioic acid was dissolved in distilled water and made up to a total ...

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Ethane, C_2H_6 , and disilane, Si_2H_6 , are both hydrides of group 4 elements with similar structures but with different chemical properties. (a) (i) Deduce the Lewis (electron dot) structure for Si_2H_6 showing all valence electrons. [1] (ii) State and explain the H-Si-H bond angle in Si_2H_6 . [2]

CHEMISTRY Candidate session number STANDARD LEVEL PAPER 2

[2] (b) Marit arranged the values she found in Table 12 into an energy cycle. $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$ $\Delta H_f^\ominus / \text{kJ mol}^{-1}$: $\text{C}_2\text{H}_4(\text{g})$: -110 , $\text{H}_2(\text{g})$: 0 , $\text{C}_2\text{H}_6(\text{g})$: -84 (hydrogenation) (l) Calculate the value for the enthalpy of hydrogenation of ethene from the energy cycle.....

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